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Canonical form in organic chemistry

It is pointed out in the comments and I repeat that the premise of the question is already fake. The problem considers not two different ions, but two mesome depictions of the same ions. This is also evident by the mesomeric arrows between the two forms. Because mesomes do not have geometric changes, both depictions are the same in all aspects and have the same (resonance-influenced) energy. Therefore, a book that claims that either of the two has low energy is wrong. What is commonly done, however, is to determine more likely resonant structures. This is, for example, when DMAP (4-N, N'-dimethyl aminopyridine), commonly used to describe the strong and stable properties of certain aromatics, is used as a carboxidin. We can try it for your structure. If you compare the two, you will see that many features are the same. The main difference is that A has a triple substitution double bond, B has secondary carbocation only as a bi-minute (and terminal) double bond B, and A has major carvocation (although both are allies). Because the aric stabilization of the cathion is much stronger than hyperconjugation, the second bullet can be discarded as essentially equal in both resonant structures. Therefore, we need to discuss with the position of the double coupling. Both double bonds can benefit from hyperconjugation twice (A primary ions cannot geometrically hyperconjugate), while in A two hyperconjugations stabilize different carbons. $\$text{mono replacement} < \text{terminal di substitution} < \text{(cis} < \text{(trans} < \text{tri substitution} < \text{tetra substitution}\$\$\$ \$\$ A is slightly better than B, so it contributes slightly towards the overall structure. As mentioned in the comments described in your book: Structure B is not a plane, so coupling is not possible in it, you would like to think that the following depiction in Figure 1 is suitable for two resonant forms: Figure 1: Under the assumption of a chair-structured cyclohexysyl structure in Figure 1: B. Different geometries are assumed for both structures B and A. And this may be valid. However, there is no way for Ccation B to adopt a chair cycloheksan configuration. Instead, as mentioned above, the following structure in Figure 2 is much more likely and shows that we are dealing with resyming: Figure 2: Depiction of mesome forms without geometrical changes. There is a very good way to optimize the geometry so that the trajectory can form an allyr system, so you can see that the discussion of non-residedity is bogus. Look forward to it! One way to represent the actual structure of molecules, including non-locale bonds, is to draw several possible structures and assume that the actual molecules are hybrids of them. These structuresIt's called a normal form, but it's not an actual structure. In other words, the molecules do not shift rapidly between them, and the given compound has a single actual structure. Its structure is always the same and is considered to be a weighted average of all normal forms. Drawing normal forms and deriving true structures from them is guided by certain rules, including: 1. All normal forms must be Bonafides Lewis structures (Sec. 1.F).For example, none of them can have carbon with five bonds. 2. The position of the nucleus must be the same for all structures. This means that when drawing different normal shapes, electrons are simply arranged differently. For this reason, the method of symultion representing resonance is easy to devise: the resonant interaction between the benzene ring and chlorine, which calls hyperconjugation (Sec.2.M, can be expressed as shown in 27 or 28, and both representations have been used in the literature to conserve space. However, the 27 curved arrow method is not used because the arrows in this document are used to represent the actual movement of electrons in reactions. Expressions such as 28 are used occasionally, but more often one or more canonical formats are used. The rule used in dashed lines, such as 28, is that joins that exist in all normal forms are drawn as solid lines, and joins that do not exist in all formats are drawn as dashed lines. The disadvantage of this model is that it is difficult to track electron transfer in reactions associated with benzene ring. For this reason, one of the normal forms is most frequently used, as described above. In most resonances, σ is not involved, only π or non-shared electrons are used. The findings mean that writing one normal form for a molecule allows other molecules to π simply by electrons that are moved and not shared. 3. All atoms participating in the resonance (i.e., covered with non-localized electrons) must lie on a plane or most of it (see Sec. 2.G.). This, of course, does not apply to atoms with the same bond in all normal forms. The maximum overlap of the p-orbits lead to planarity. 4. All normal forms must have the same number of unsynthes electrons. Therefore, the diradical structure •CH2CH=CH-CH2• is not an effective normal form of butadiene. 5. Since the actual molecular energy is lower than in any form, the de-localization phenomenon is a stabilization phenomenon. All normal forms do not contribute equally to true molecules. Each form benefits proportionally to stability, with the most stable form contributing the most. Therefore, in the case of ethylene, the format has a very high energy compared to CH2 = CH2 and essentially does not contribute at all. This argument was applied to contribute equally to butadiene .39 equivalent normal forms (for example, 1 and 2). The greater the number of important structures that can be written, the larger theResonant energy, others equal. It is not always easy to determine the relative stability of a fictional structure. Chemists are often guided by intuition. Structures with more cojoins are usually more stable than fewer structures (cf.6 and 7). B. Stability is reduced by an increase in charge separation. Structures with formal charges are less stable than unchar charged structures. A structure with two or more formal charges is usually almost no contribution. A particularly unfavorable structure is one with two similar charges for adjacent atoms. Structures that carry negative charges to more electrically negative atoms are more stable than those in electrically negative atoms with less charge. Therefore, the enorate anion 30 is more stable than the normal form 29. Similarly, positive charges are best carried against low electrically negative atoms. d. Structures with distorted bond angles or lengths are unstable (e.g., etang contributor 31). A glossary of back-content forward organic chemical acetate ion resonance hybrids is shown. (d) The reticulity of the alkyll haligenate can be inferred by reacting the substance with Ag salt (AgNO3). Based on the binding dissoional energy RI, it is the maximum rescivity due to the maximum size of the I atom and the minimum binding dissoional energy. Free radical substitution reactions These types of reactions are given by alkanes, alkenes and some aromatic compounds in the presence of sunlight or high temperatures. For free radical substitution in Alken, search for agaryl position and replacement at this location for intermediate stabilization by co-ed. In toluene and its derivatives, the free radical substitution reaction occurs at the benzyl position. If there is a possibility of replacing both agrical and benzyl positions together than substitution at the benzyl level, preferences due to more intermediate stability are given. [3] The different contributing structures of resonant molecules differ only in the distribution of electrons, not in the arrangement of atoms, i.e. the position of atoms does not change (different from tatmarism). In addition, each normal form must have the same number of vs. electrons, and its position may be different. Usually, there is a large number of normal forms in the compound, and its resonant energy is large, and stability is high. (i) Differences in the stability of various normal forms of compounds : Not all normal forms of compounds are equally stable. One form may be more stable than the other, thus contributing to the actual resonant hybrid. The main factors governing the stability of the positive pole structure are: 1. The non-polar structure is more stable than the polar structure. 2. The more is the number of shared joins in the normal structure, and the more its stability. 3. In a polar structure, its structure is negative and positiveThey exist in the most electrically negative atoms and the most potential atoms, respectively. 4. Of the two normal structures, those with finished octets of various atoms (or hydrogen double) are more stable, even when more electrically negative atoms have a positive charge. 5. A normal structure with positively charged electron-deficited atoms is much less stable. 6. Also, structures where the charge is on atoms close to each other are the least stable. (ii) Rules for writing resonant structures a) No existence: Resonant structures exist only on paper. Resonant structures are useful because they can describe molecules, radicals, and ions with insufficient single Lewis structures. We write two or more Lewis structures, calling them the resonant structures of resonant contributors. Connect these structures with double-head arrows and say that real molecules, radicals and ions are hybrids of them all. (b) When writing a resonant structure, the position of the nucleus must be the same in all resonant structures© all structures must have a proper Lewis structure (d) charge separation is required, so it is a structure with greater energy to separate charge energy and therefore less stable. (iii) Re-acting energy The difference in energy between hybrids and the most stable normal structures is called re-acting energy. Any halogen or electrical negative element prefers only an O with a negative charge for stability, in the resyming structure of acrylic aldehydes, which always prefers negative charges when there are options for both +ve and –ve charge. There are always two types of coupling, one localized and the other decentralized, responsible for resoning the (contributing) structure. Structural isomers are real molecules in which atoms are linked in different ways to form different skeletons, but the resonant structure is not real. They are written when a single electronic structure cannot properly represent the actual structure.

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